

Glycol - Purpald®-Periodate Method

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Applications and Industries

Airplane de-icing operations, industrial process waters, cooling water systems, stormwater, potable water

NOT applicable for seawater analysis

References

Purpald® developed by Aldrich Chemical Company

Fritz, James S. and Schenk, George H., Quantitative Analytical Chemistry, 4th ed., p. 277 (1979)

Chemistry

Periodic acid oxidizes ethylene glycol and propylene glycol to formaldehyde. In a highly alkaline solution, and in conjunction with the oxidizing agent persulfate, formaldehyde reacts with Purpald to form a purple-colored complex. Test results are expressed as ppm (mg/L) ethylene glycol. To convert results to ppm propylene glycol, multiply by 2.

Note: This chemistry does not differentiate between ethylene and propylene glycol.

Available Analysis Systems

Visual colorimetric: CHEMets®

Storage Requirements

Products should be stored in the dark and at room temperature. High temperatures can cause the ampoule reagent to expire prematurely. Glass-like crystals in an unused ampoule are an indication of reagent deterioration.

Safety Information

Safety Data Sheets (SDS) are available upon request and at www.sdsfetch.com. Read SDS before using these products. Breaking the tip of an ampoule in air rather than water may cause the glass ampoule to shatter. Wear safety glasses and protective gloves.

Accuracy Statement

Statements of accuracy are based on laboratory tests performed under ideal testing conditions using standards of known concentration prepared in deionized water.

CHEMets® kit: ± 1 color standard increment

Interference Information

- This test procedure is somewhat temperature dependent. Extremely high or low temperatures may affect the rate of the reaction, causing erroneous results. For best results, samples should be less than 40°C.
- Strong oxidizers may cause false positive results, and strong reducing agents may cause low test results.
- The chemical reaction occurs at a high pH. Sample pHs below 4 or samples buffered to a low pH may not develop the proper color.
- Formaldehyde reads positively with this chemistry if present at any level. Other aldehydes may also cause high test results or may develop a different color with the reagent.
- Ferrous and ferric iron may cause off-color test results, making a visual color match difficult or impossible.
- Ethanol up to at least 1% does not interfere.
- Methanol at percent levels causes a false positive result. Methanol-based solutions also may cause formation of a precipitate during analysis and may cause an off-color to develop in the test ampoule.
- Diethylene and triethylene glycol are not measured with this chemistry.
- Samples with high dissolved solids content may cause the reagent to precipitate.
- This chemistry is not applicable for the analysis of seawater samples.
- Sample color or turbidity may make a color match difficult.